

## ALUMINIO E HIDRÓXIDO SÓDICO

### Hidróxido sódico 2N y polvo de aluminio

Trabajando como en prácticas anteriores, con un poco de aluminio en polvo y una gota de hidróxido sódico 2N(fig.1), que reacciona en frío, lentamente al inicio (fig. 2, y 3), con una pequeña producción de hidrógeno. A los 10 minutos, la reacción se hace mas violenta, atacando el aluminio, con desprendimiento de hidrógeno de forma espectacular (fig.4-5-6). Las burbujas se van haciendo mas grandes, (fig 7-15), formándose aluminato sódico.



Fig.1

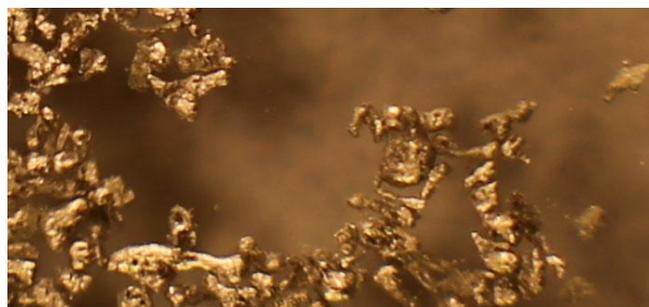


Fig.2 (detalle)



Fig.3(detalle)



Fig.5

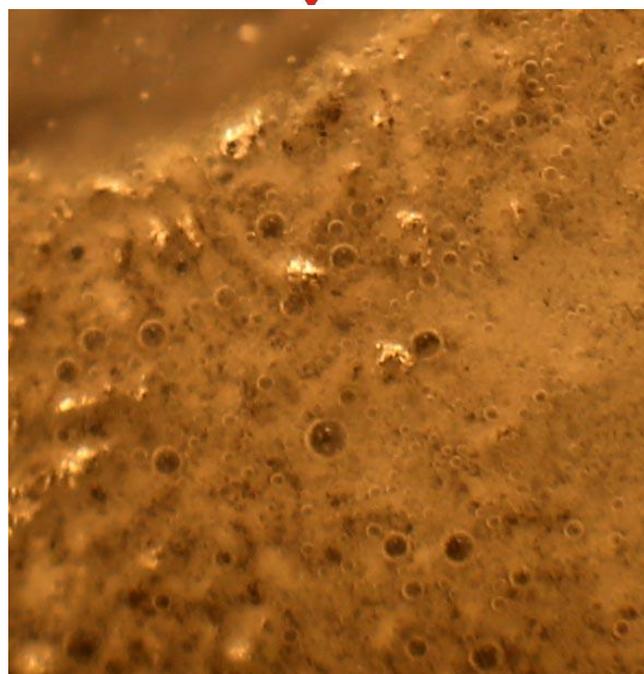


Fig.4

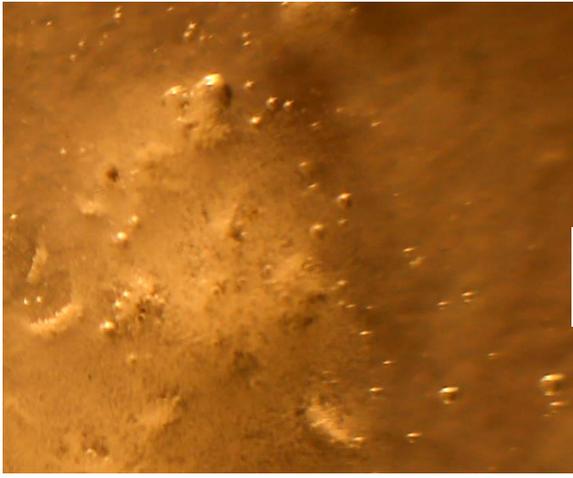


Fig.6(detalle)

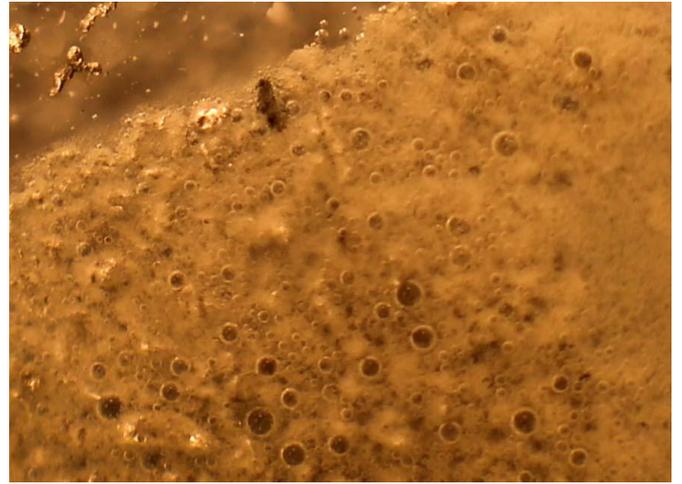
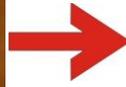


Fig7

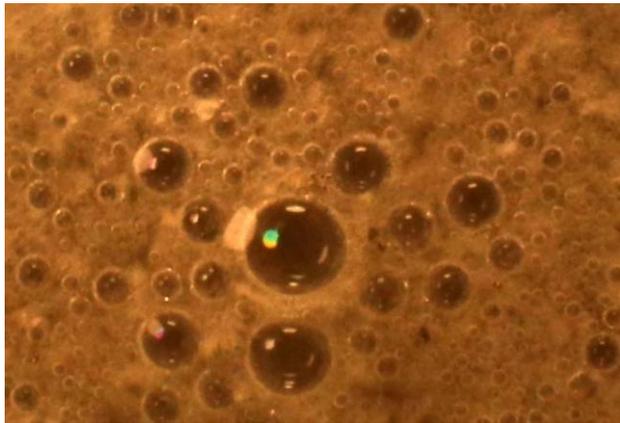


Fig.9 (detalle)

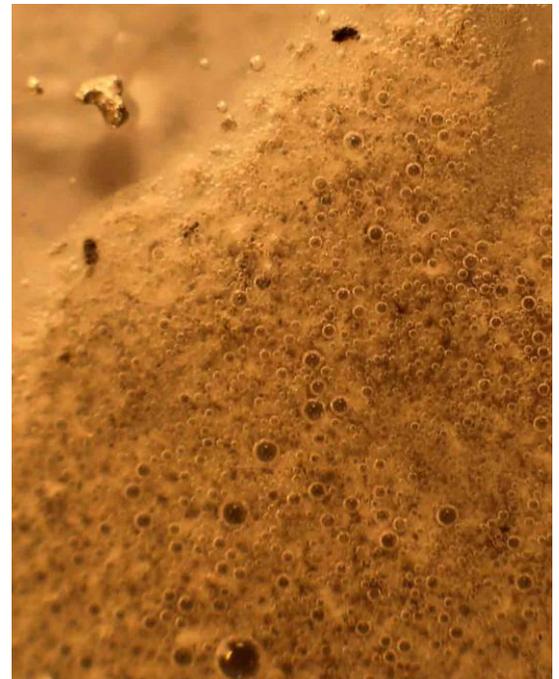


Fig.8

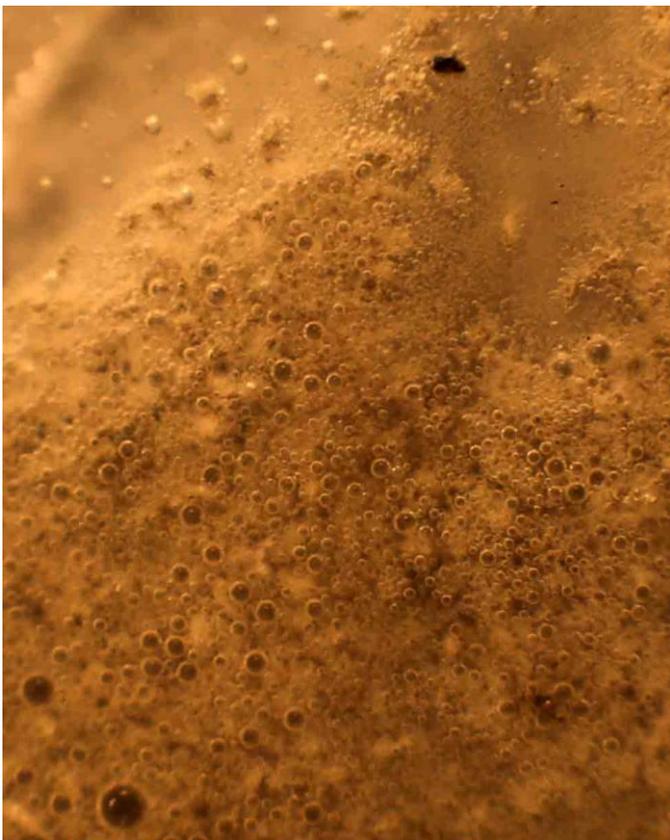
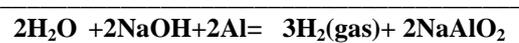
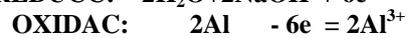
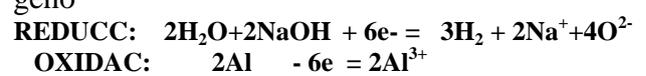


Fig.10

Las reacciones químicas que se producen son redox en medio básico ocasionando aluminato sódico e hidrógeno



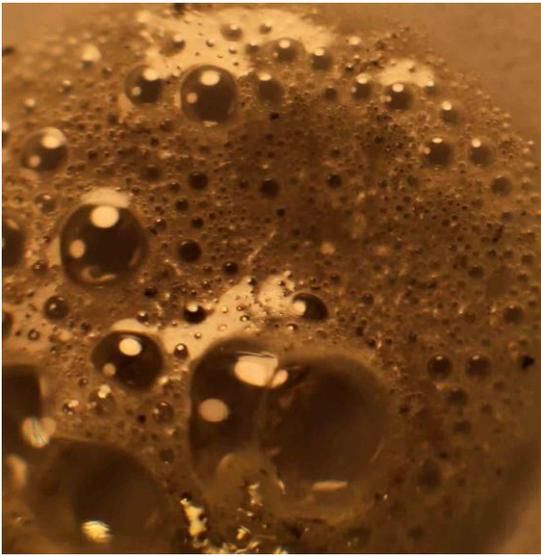


Fig.11



Fig.12

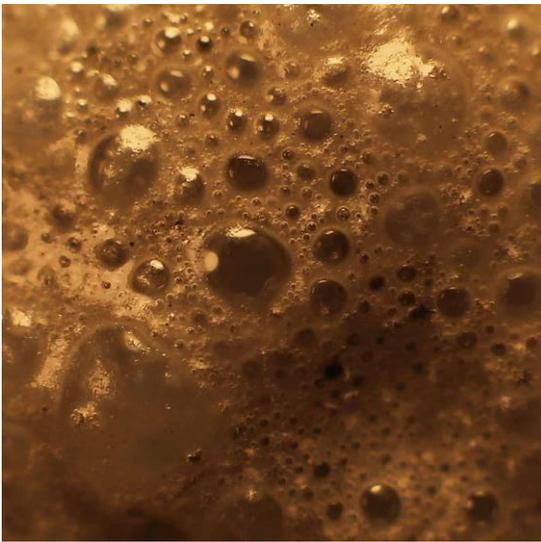
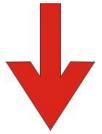


Fig.14



Fig.13

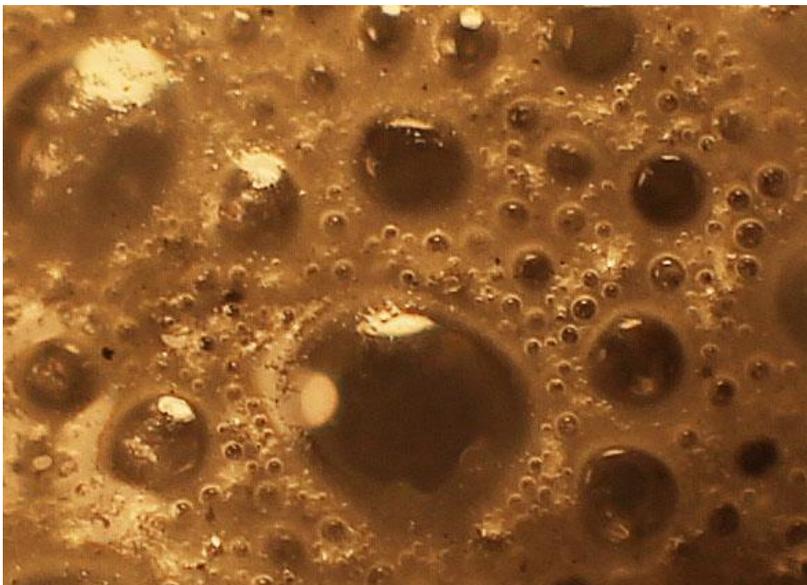


Fig.15